

Chapter 8 Chemical Equations Reactions Answer Key

~~CHAPTER 8 Chemical Equations and Reactions Chapter 8 - Chemical Equations and Reactions~~
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~~CHAPTER 8 Chemical Equations and Reactions~~

Chemical Equations and Reactions The evolution of energy as light and heat is an indication that a chemical reaction is taking place. CHAPTER 8
Fireworks

~~Chapter 8—Chemical Equations and Reactions~~

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chemical reaction in which one element replaces another element in a compound double-displacement reaction a reaction in which a gas, a solid
precipitate, or a molecular compound forms from the exchange of ions between two compounds

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CHAPTER 8 Chemical Equations and Reactions. SECTION 1 Describing Chemical Reactions. OBJECTIVES. 1. List three observations that suggest that a
chemical reaction has taken place. 2. List three requirements for a correctly written chemical equation. 3. Write a word equation for a correctly
written chemical equation.

~~Chapter 8 Chemical equations and reactions(no example ...~~

Chapter 8: Chemical Equations and Reactions. Doing so actually changes the chemical formula, and that does not happen. Adjust the number of
atoms on each side of the equation by changing the coefficients, the larger number in front of the compound. Check to make sure the number and
type of each atom is the same on both sides of the equation.

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~~8-Chemical Equations and Reactions~~

formula equation for a given chemical reaction. Balance a formula equation by inspection. FIGURE 8-1 The decomposition of ammonium dichromate proceeds rapidly, releasing energy in the form of light and heat.

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Chapter 8 - Chemical Equations & Reactions. We will conclude the chapter by learning about the different types of chemical reactions; students will be able to differentiate between synthesis, decomposition, single replacement, and double replacement reactions and be able to predict the products (if any) of a given reaction.

~~CHAPTER 8 Chemical Equations and Reactions~~

Chapter 8 - Chemical Equations and Reactions. Describing Chemical Reactions _____ - process in which one or more substances are changed into one or more different substances. Reactants - original substance. Products - resulting substance. Described by _____ Chemical equation - represents, with symbols and formulas, the identities and ...

~~Chapter 8 Chemical Equations~~

CHAPTER 8 REVIEW Chemical Equations and Reactions SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. Match the equation type on the left to its representation on the right. c synthesis (a) $AX + BY \rightarrow AY + BX$ d decomposition (b) $A + BX \rightarrow AX + B$ b single-displacement (c) $A + B \rightarrow AX$ a double-displacement (d) $AX \rightarrow A + X$

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bc any instructions shown in a chemical equation can help you be sure the reaction turns out the way it should. when info about conditions of a reaction is desired, the _____ shows it! arrow (ex: $N_2(g) + 3H_2(g) \xrightarrow{\text{above arrow reads 350 degrees celcius, 2500 kPa}} 2NH_3(g)$)

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Chapter 8: Chemical Equations and Reactions 1. The Equation must represent known facts. 2. The Equation must contain the correct formulas for the reactants and products.

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8.1 The chemical equation reactants - the substances entering the reaction products - the substances formed In a chemical reaction atoms are neither created nor destroyed. All atoms present in the reactants must also be present in the products ex. thermite reaction 1. $Al + Fe_2O_3 \rightarrow Fe + Al_2O_3$ reactants products 2. $2Al + Fe_2O_3 \rightarrow 2Fe + Al_2O_3$

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Questions for Class 10 ...

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A word equation and a formula equation are both qualitative. Although they both show the reactants and products of a chemical reaction, they give no information about the amounts of reactants or products. A chemical equation, however, does show the relative amounts of the reactants and products in a chemical reaction.

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1. the coefficients of a chemical reaction indicate relative, not absolute, amounts of reactants and products
2. the relative masses of the reactants and products of a chemical reaction can be determined from the reaction's coefficients
3. the reverse reaction for a chemical equation has the same relative amounts of substances as the forward ...

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The usual evidence of an reaction is a solid precipitant, the evolution of a gas, or the formation of water. (Solubility rules). Neutralization is the reaction b/t an acid (H⁺ cation) and a base (OH⁻ anion) to form an ionic salt and water.

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