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The maximum number of electrons an orbital can contain is 2 electrons.
Describe the relative orientations of the orbitals related to an atom's 2p sublevel.
The relative orientations of the orbitals in an atom's 2p sublevel is 3 orbitals, 2p_x, 2p_y and 2p_z that look like figure eights and are perpendicular to each other

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an electron can only... Einstein
Considered light to be both a wave and a
particle of energy.

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states that a maximum of two electrons
can occupy a single atomic orbital, but if
only if the electrons have opposite spins
states that each electron occupies the
lowest energy orbital available states
that single electrons electrons with the
same spin must occupy each-energy
orbital before additional electrons with
opposite spins can occupy the ...

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This video describes light as a particle
and wave. It also describes matter and
quantum of energy.

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The Atom. A. Represented by v (ν). This

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is the number of waves that pass a given point per second. The SI unit for frequency. The speed of light is an... The wave's height from the origin to the crest. Which is the correct formula for the speed of light. None of the above.

Science / Chapter 5 - electrons in atoms (handouts)

Chapter 5 - Electrons in Atoms - 5.1
Revising the Atomic Model - 5.1 Lesson
Check - Page 132: 5 Answer Quantized energies means that electrons are moved between energy levels by gaining or losing a certain amount of energy.

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Schiller. STUDY. Flashcards. Learn.
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Created by. jenkins6. Electrons in atoms.
Terms in this set (18) energy levels.
fixed energies an electron can have
(circles around nucleus) quantum.

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Chemistry chapter 5 electrons in an atom. Orbitals of equal energy or each occupied by one electron before any or both of you bye-bye second electron and that each of the single electrons must have the same spin If you fill 1s and 2s then you get to 2p which have 3 boxes you have to fill each box with one electron before you fill the first of the three boxes with a second electron.

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CHEMISTRY CHAPTER 5 NOTES 5.1 -
Light and Quantized Energy. • The
Nuclear Atom and Unanswered

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Questions o Although Rutherford's scientific model of an atom was a breakthrough, it lacked detail about how electrons occupy the space surrounding the nucleus of an atom.

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A three-dimensional region around the nucleus of an atom that... electrons occupy orbitals of the lowest energy levels first. an arrangement of electrons around the nucleus of an atom (loo.... The amount of energy an electron can have.

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